

Candidate 9 evidence

Vitamin C content in fruit Juice.

Aim: to calculate the mass of Vitamin C in different flavours of ~~ocean~~ ocean spray cranberry juice per 100ml

~~Vitamin~~ Underlying Chemistry

Vitamin C, also ~~known~~ known as ascorbic acid, facilitates lots of repair processes and metabolic reactions within the body. It isn't stored very well in the body and is destroyed by heat. It occurs in fresh fruit in a great abundance and also in most fresh vegetables. Ascorbic acid is also a very good reducing agent.

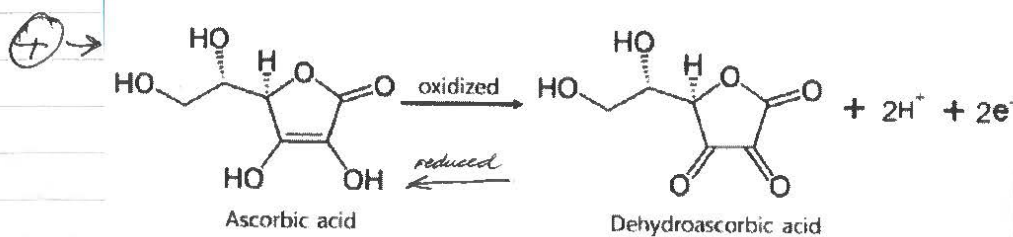
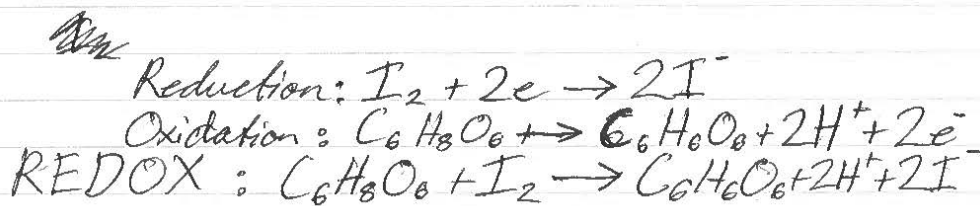
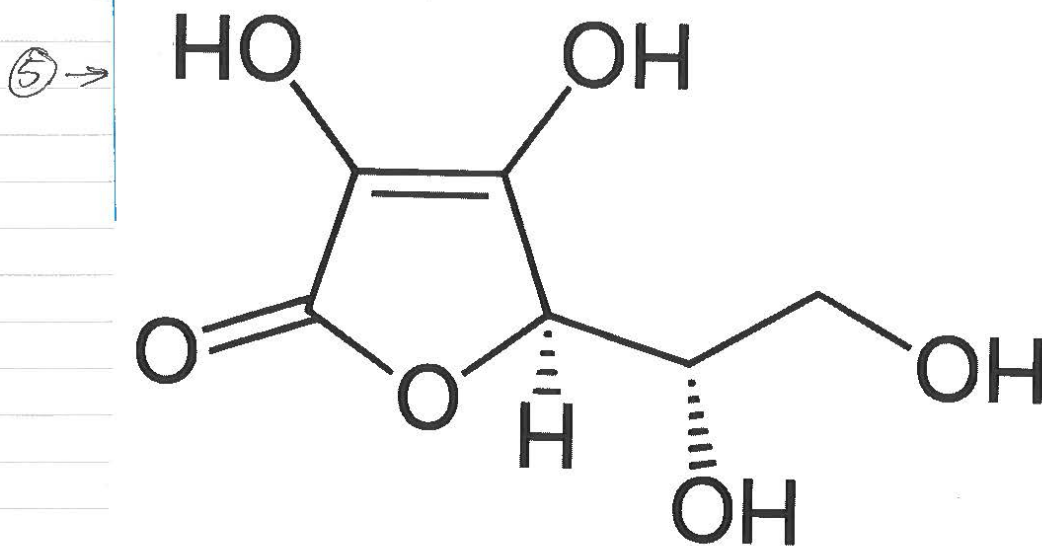
In a REDOX titration, there is a transfer of electrons between the reactants. A reducing agent is titrated with an oxidising agent, (In this case it ~~is~~ is Vitamin C and Iodine) or vice versa.

~~When the transfer of the electrons of the~~
When the reducing agent loses its electrons and donates them to the oxidising agent that then accepts them, one of the reactants is oxidised and the other one is reduced.

Vitamin C (ascorbic acid) is a reducing agent due to the loss of the hydrogen atoms from the OH's, which makes Vitamin C a hydrogen donor, and thus, a reducing agent.

1.

Full chemical structure of Vitamin C:



2.

Data Collection + Handling

The experiment that we did consisted of titrating ~~the~~ $0.0025 \text{ mol l}^{-1}$ Iodine solution into different juice flavours until a colour change appeared due to the starch indicator that we ^{added}.

~~Approx~~ The apparatus we used consisted of:

- a burette
- a white tile
- a pipette
- a burette clamp stand
- a conical flask
- a filter funnel
- a dropper
- beakers containing our reactants
- a ~~pipette~~ ^{Pipette} filler

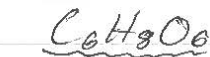
We also wore gloves to prevent the iodine from staining our skin.

Classic Cranberry 25 cm³

	Initial Volume cm ³	Final Volume cm ³	Volume Added cm ³
Rough	9.2	22	12.8
1	22	32.4	10.4
2	32.4	42.4	10.0
3	0	10.5	10.5

Concordant

10.45 = avg titre volume



$$n = 0.0001045 \text{ mol}$$

$$n = 0.0001045 \text{ mol}$$

$$c = 0.00418 \text{ mol l}^{-1}$$

$$c = 0.0025 \text{ mol l}^{-1}$$

$$v = \frac{25}{1000} = 0.025 \text{ L}$$

$$v = \frac{10.45}{1000} = 0.01045 \text{ L}$$

mol Ratio

$$1 : 1$$

$$\frac{0.0001045}{0.0001045} = 1$$

G_{Fm} of C₆H₈O₆ = 176

$$m = n \times F_m$$

$$c = \frac{n}{v}$$

$$n = c \times v$$

$$m = 0.0001045 \times 176$$

$$= \frac{0.0001045}{0.025}$$

$$= 0.0025 \times 0.01045$$

$$= 0.018392 \text{ g}$$

$$= 0.00418 \text{ mol l}^{-1}$$

$$= 0.00026125 \times 4 = 0.001045 \text{ mol}$$

$$= 18.39 \text{ mg/100 ml}$$

$$= 4.18 \text{ mg/100 ml}$$


3.

Apple Cranberry 25 cm³

	Initial Volume cm ³	Final Volume cm ³	Volume Added cm ³
Rough	0	19.6	19.6
1	19.6	39.4	19.8
2	0	19.8	19.8
3			

19.8 concordant
19.8 avg titre = 19.8

C₆H₈O₆



$n = 0.000198 \text{ mol}$
 $C = 0.00792 \text{ mol L}^{-1}$
 $V = \frac{25}{1000} = 0.025 \text{ L}$

I₂

$n = 0.000198 \text{ mol}$
 $C = 0.0025 \text{ mol L}^{-1}$
 $V = \frac{19.8}{1000} = 0.0198 \text{ L}$

mol Ratio
1 : 1 $n = c \times v$

$C_{6}H_{8}O_{6}$ of $C_{6}H_{8}O_{6} = 176$

$C = \frac{n}{V}$
 $= \frac{0.000198}{0.025}$
 $= 0.00792 \text{ mol L}^{-1}$

$m = n \times F_m$
 $m = 0.000198 \times 176$
 $m = 0.034848 \text{ g}$
 $\times 1000$
 $\therefore = \underline{34.8 \text{ mg/100 ml}}$

$0.000198 \times 0.000198 = 0.0025 \times 0.0198$
 $= 0.0000495 \text{ mol} \times 4$

$C_{6}H_{8}O_{6}$ of $I_{2} = 176$ $= 0.000198 \text{ mol}$


4.

Raspberry Cranberry 50 cm³

	Initial Volume cm ³	Final Volume cm ³	Volume Added cm ³
Rough	0	10.5	10.5
1	10.5	20.8	10.3
2	20.8	31.0	10.2
3			

concordant
avg titre
= 10.25

C₆H₈O₆



$n = 0.00005125 \text{ mol}$
 $c = 0.001025 \text{ mol L}^{-1}$
 $V = \frac{50}{1000} = 0.05 \text{ L}$

I₂

$n = 0.00005125 \text{ mol}$
 $c = 0.0025 \text{ mol L}^{-1}$
 $V = \frac{10.25}{1000} = 0.01025 \text{ L}$

C₆H₈O₆ at

$C = \frac{n}{V}$
 $= \frac{0.00005125}{0.05}$
 ~~$= 0.000001025$~~
 $= 0.001025 \text{ mol L}^{-1}$

$m = n \times F_m$
 $= 0.00005125 \times 176$
 $= 0.00902 \text{ g}$
 $\times 1000$
 $\therefore = 9.02 \text{ mg} / 100 \text{ ml}^*$

mol Ratio

1 : 1

$0.00005125 : 0.00005125$

$n = c \times V$
 $= 0.0025 \times 0.01025$
 $= 0.000025625 \text{ mol} \times 2$
 $= 0.00005125 \text{ mol}$

Internet Sources:

① → CRANBERRY CLASSIC OCEAN SPRAY:

Nutritional Information

Typical Values	per 100ml (%RI**)
Energy	86kJ/20kcal
Carbohydrate	4.3g
Of which sugars	4.3g
Vitamin C	24mg (30%)
**Reference Intake of an average adult (8400kJ/2000kcal)	-
Contains negligible amounts of Fat, Saturates and Protein and Salt	-

② → CRANBERRY AND RASPBERRY OCEAN SPRAY:

Nutritional Information

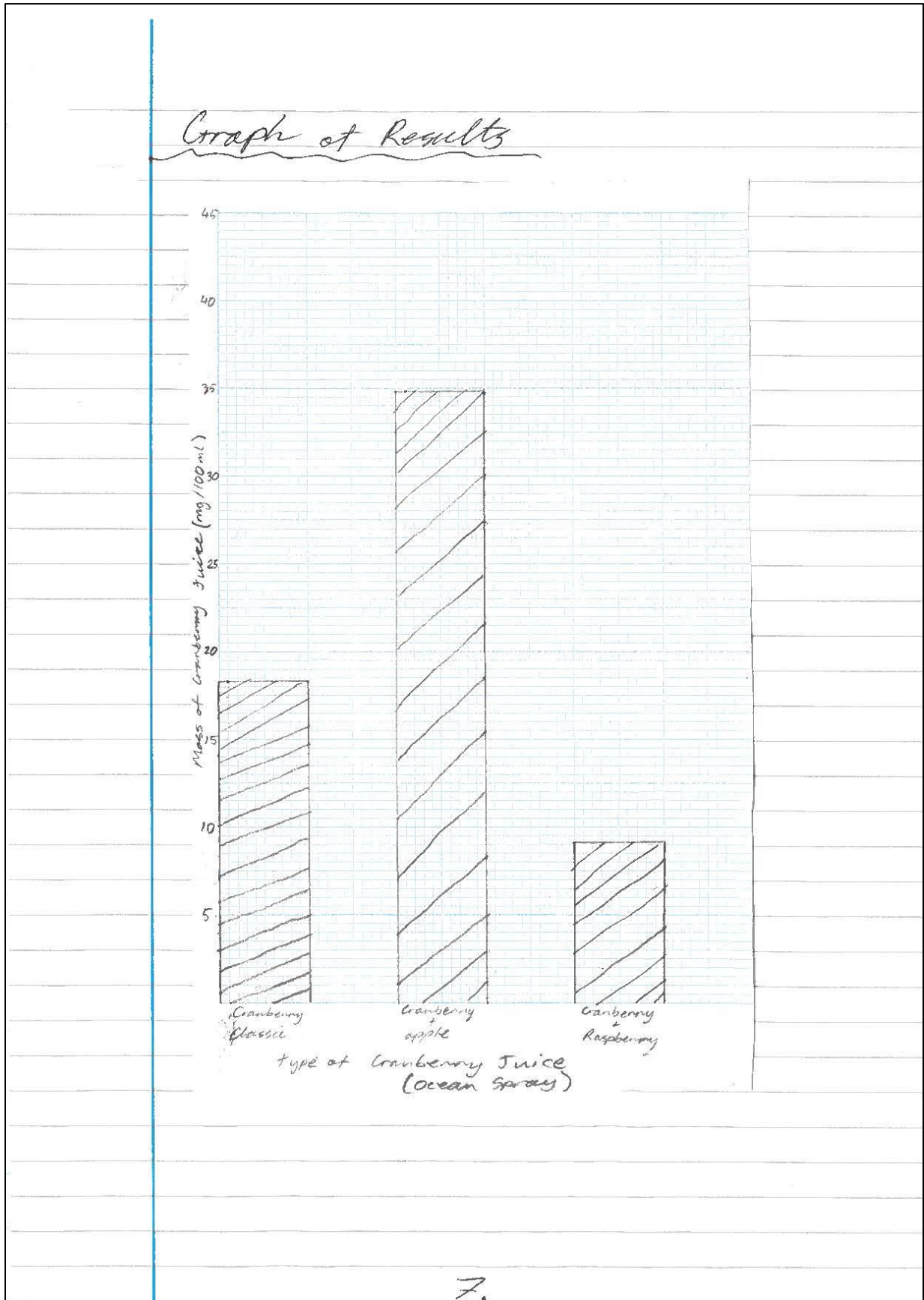
Typical Values	per 100 ml (%RI**)
Energy	89 kJ / 21 kcal
Carbohydrate	4.9 g
of which Sugars	4.5 g
Salt	0.03 g
Vitamin C	12 mg (15%)
**Reference Intake of an average adult (8400kJ/2000kcal)	-
Contains negligible amounts of Fat, Saturates and Protein	-

③ → CRANBERRY AND APPLE OCEAN SPRAY:

Nutritional Information

Typical Values	per 100ml	(%RI**)
Energy	87 kJ	
-	20 kcal	
Carbohydrate	4.8 g	
of which Sugars	4.1 g	
Vitamin C	20 mg	(25%)
Contains negligible amounts of Fat, Saturates, Protein and Salt	-	-
**Reference Intake of an average adult (8400kJ/2000kcal)	-	-

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<u>Analysis</u>		
Type of Juice	My Results <u>mg</u>	Internet Source <u>mg</u>
Cranberry Classic	18.39	24
Cranberry + Apple	34.8	20
Cranberry + Raspberry	9.02	12

~~On~~ ^{On} comparing my results to the internet source, it shows no general trend ~~and~~ in results as my highest result, (Cranberry + Apple) is not the highest according to the internet, which shows Classic cranberry as the highest.

~~Conclusion~~ Conclusion

The data in my experiment has been calculated to find the mass of Vitamin C in different flavours of ocean spray cranberry juice, we got ~~the~~ ^{a set of} results but these do not match my trends according to the internet.

Evaluation

- The ~~data~~ data from the internet is from ~~a~~ a website ^{and follows} U.K. food regulations law's so it has to be completely accurate.
- In one of our experiments we had to increase the volume of juice, (Cranberry + Raspberry) to

8.

50 cm³ instead of 25 cm³ as the colour change was happening too quickly so our results were ~~flawed~~ ^{inconsistent}. This change allowed us to get more accurate results that were closer to the ones from the internet source.

* Information Links / References

① → <https://www.tesco.com/groceries/en-GB/products/254859384>

DATE ACCESSED 23/11/2023

② → <https://www.tesco.com/groceries/en-GB/products/257630874>

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③ → <https://www.tesco.com/groceries/en-GB/products/304384270>

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④ → https://www.researchgate.net/figure/Chemical-structure-of-ascorbic-acid-and-dehydroascorbic-acid-as-depicted-through-a-redox_fig1_312229647

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⑤ → https://en.wikipedia.org/wiki/Vitamin_C

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9.